

Methyl 2-(3-benzoylthioureido)acetate

Ibrahim N. Hassan, Bohari M. Yamin and Mohammad B. Kassim*

School of Chemical Sciences and Food Technology, Faculty of Science and Technology, Universiti Kebangsaan Malaysia, UKM 43600 Bangi Selangor, Malaysia
Correspondence e-mail: mbkassim@ukm.my

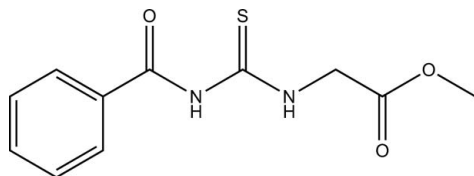
Received 22 October 2009; accepted 3 November 2009

Key indicators: single-crystal X-ray study; $T = 298$ K; mean $\sigma(\text{C}-\text{C}) = 0.006$ Å; R factor = 0.089; wR factor = 0.172; data-to-parameter ratio = 14.7.

In the title compound, $\text{C}_{11}\text{H}_{12}\text{N}_2\text{O}_3\text{S}$, the methyl acetate and benzoyl groups adopt a *cis-trans* configuration with respect to the thiono S atom across the C–N bonds. An intramolecular N–H···O hydrogen bond is observed. In the crystal packing, molecules are linked by intermolecular N–H···S and C–H···O hydrogen bonds to form a two-dimensional network lying parallel to (101).

Related literature

For bond-length data, see: Allen *et al.* (1987). For related structures, see: Hassan *et al.* (2008*a,b,c*); Yamin & Hassan (2004). For the preparation, see: Hassan *et al.* (2008*a*).



Experimental

Crystal data

 $\text{C}_{11}\text{H}_{12}\text{N}_2\text{O}_3\text{S}$
 $M_r = 252.29$

 Monoclinic, $P2_1/n$
 $a = 14.5804$ (15) Å

 $b = 4.9740$ (5) Å

 $c = 16.9133$ (16) Å

 $\beta = 96.210$ (2)°

 $V = 1219.4$ (2) Å³
 $Z = 4$

 Mo $K\alpha$ radiation

 $\mu = 0.26$ mm⁻¹
 $T = 298$ K

 $0.48 \times 0.14 \times 0.06$ mm

Data collection

Bruker SMART APEX CCD area-detector diffractometer

Absorption correction: multi-scan (SADABS; Bruker, 2000)

 $T_{\min} = 0.884$, $T_{\max} = 0.984$

6881 measured reflections

2264 independent reflections

 1600 reflections with $I > 2\sigma(I)$
 $R_{\text{int}} = 0.048$

Refinement

 $R[F^2 > 2\sigma(F^2)] = 0.089$
 $wR(F^2) = 0.172$
 $S = 1.24$

2264 reflections

154 parameters

H-atom parameters constrained

 $\Delta\rho_{\text{max}} = 0.26$ e Å⁻³
 $\Delta\rho_{\text{min}} = -0.18$ e Å⁻³
Table 1

Hydrogen-bond geometry (Å, °).

$D-H\cdots A$	$D-H$	$H\cdots A$	$D\cdots A$	$D-H\cdots A$
$\text{N2}-\text{H2A}\cdots\text{O1}$	0.86	2.02	2.676 (4)	132
$\text{N1}-\text{H1A}\cdots\text{S1}^{\text{i}}$	0.86	2.77	3.547 (3)	151
$\text{C9}-\text{H9A}\cdots\text{O2}^{\text{ii}}$	0.97	2.54	3.358 (6)	142
$\text{C9}-\text{H9A}\cdots\text{O2}^{\text{iii}}$	0.97	2.58	3.211 (5)	123

 Symmetry codes: (i) $-x + 1, -y, -z + 1$; (ii) $x, y + 1, z$; (iii) $-x + \frac{1}{2}, y + \frac{1}{2}, -z + \frac{3}{2}$.

Data collection: SMART (Bruker, 2000); cell refinement: SAINT (Bruker, 2000); data reduction: SAINT; program(s) used to solve structure: SHELXS97 (Sheldrick, 2008); program(s) used to refine structure: SHELXL97 (Sheldrick, 2008); molecular graphics: SHELXTL (Sheldrick, 2008); software used to prepare material for publication: SHELXTL, PARST (Nardelli, 1995) and PLATON (Spek, 2009).

The authors thank Universiti Kebangsaan Malaysia for providing the facilities and the Ministry of Science, Technology and Innovation for the research fund (Nos. UKM-ST-01FRGS0016–2006, UKM-GUP-BTT-07-30-190 and UKM-OUP-TK-16-73/2009).

Supplementary data and figures for this paper are available from the IUCr electronic archives (Reference: CI2952).

References

- Allen, F. H., Kennard, O., Watson, D. G., Brammer, L., Orpen, A. G. & Taylor, R. (1987). *J. Chem. Soc. Perkin Trans. 2*, pp. S1–19.
- Bruker (2000). SMART, SAINT and SADABS. Bruker AXS Inc., Madison, Wisconsin, USA.
- Hassan, I. N., Yamin, B. M. & Kassim, M. B. (2008*a*). *Acta Cryst.* **E64**, o1727.
- Hassan, I. N., Yamin, B. M. & Kassim, M. B. (2008*b*). *Acta Cryst.* **E64**, o2083.
- Hassan, I. N., Yamin, B. M. & Kassim, M. B. (2008*c*). *Acta Cryst.* **E64**, o2167.
- Nardelli, M. (1995). *J. Appl. Cryst.* **28**, 659.
- Sheldrick, G. M. (2008). *Acta Cryst.* **A64**, 112–122.
- Spek, A. L. (2009). *Acta Cryst.* **D65**, 148–155.
- Yamin, B. M. & Hassan, I. N. (2004). *Acta Cryst.* **E60**, o2513–o2514.

supplementary materials

Acta Cryst. (2009). E65, o3078 [doi:10.1107/S1600536809046169]

Methyl 2-(3-benzoylthioureido)acetate

I. N. Hassan, B. M. Yamin and M. B. Kassim

Comment

The title compound, (I), is a methyl ester derivative of glycine thiourea analogous to our previously reported ethyl-2-(3-benzoylthioureido)acetate (II) (Hassan *et al.*, 2008*a*), propyl-2-(3-benzoylthioureido)acetate (III) (Hassan *et al.*, 2008*b*) and butyl-2-(3-benzoylthioureido)acetate (IV) (Hassan *et al.*, 2008*c*). The methyl acetate fragment and the benzoyl group adopt a *cis-trans* configuration, respectively, with respect to the thiono S atom across the C—N bonds (Fig 1). The dihedral angle between the phenyl ring (C1—C6) and the central fragment (S1/N1/N2/C8/C9) is 20.12 (19)°. The bond lengths (Allen *et al.*, 1987) and angles in the molecule are in normal ranges and comparable to those of (II), (III) and (IV). The methyl acetate group (O2/O3/C9/C10/C11) is planar, with a maximum deviation of 0.023 (3) Å for atom O3. The dihedral angle between the phenyl ring and the methyl acetate group is 73.4 (2)°. An intramolecular N2—H2A···O1 hydrogen bond (Table 1) forms a pseudo-five-membered N2/H2A/O1/C7/N1/C8 ring.

Intermolecular N1—H1A···S1 and C9—H9A···O2 hydrogen bonds (Table 1) link the molecules into a two-dimensional network parallel to the (101) (Fig 2).

Experimental

The title compound was synthesized according to a previously reported method (Hassan *et al.*, 2008*a*). Yellowish crystals, suitable for X-ray analysis, were obtained by slow evaporation of a CH₂Cl₂ solution at room temperature (yield 73%).

Refinement

H atoms were positioned geometrically [N—H = 0.86 Å and C—H = 0.93–0.97 Å] and allowed to ride on their parent atoms, with $U_{\text{iso}}(\text{H}) = 1.2U_{\text{eq}}(\text{C}, \text{N})$ and $1.5U_{\text{eq}}(\text{C}_{\text{methyl}})$.

Figures

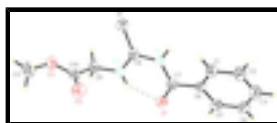


Fig. 1. The molecular structure of the title compound, with displacement ellipsoids are drawn at the 50% probability level. The dashed line indicates a hydrogen bond.

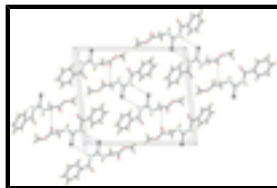


Fig. 2. Crystal packing of the title compound, viewed normal to the (101). Hydrogen bonds are shown as dashed lines.

Methyl 2-(3-benzoylthioureido)acetate

Crystal data

$C_{11}H_{12}N_2O_3S$	$F_{000} = 528$
$M_r = 252.29$	$D_x = 1.374 \text{ Mg m}^{-3}$
Monoclinic, $P2_1/n$	Mo $K\alpha$ radiation, $\lambda = 0.71073 \text{ \AA}$
Hall symbol: -P 2yn	Cell parameters from 871 reflections
$a = 14.5804 (15) \text{ \AA}$	$\theta = 1.8\text{--}25.5^\circ$
$b = 4.9740 (5) \text{ \AA}$	$\mu = 0.26 \text{ mm}^{-1}$
$c = 16.9133 (16) \text{ \AA}$	$T = 298 \text{ K}$
$\beta = 96.210 (2)^\circ$	Needle, colourless
$V = 1219.4 (2) \text{ \AA}^3$	$0.48 \times 0.14 \times 0.06 \text{ mm}$
$Z = 4$	

Data collection

Bruker SMART APEX CCD area-detector diffractometer	2264 independent reflections
Radiation source: fine-focus sealed tube	1600 reflections with $I > 2\sigma(I)$
Monochromator: graphite	$R_{\text{int}} = 0.048$
$T = 298 \text{ K}$	$\theta_{\text{max}} = 25.5^\circ$
ω scans	$\theta_{\text{min}} = 1.8^\circ$
Absorption correction: multi-scan (SADABS; Bruker, 2000)	$h = -17 \rightarrow 17$
$T_{\text{min}} = 0.884$, $T_{\text{max}} = 0.984$	$k = -5 \rightarrow 6$
6881 measured reflections	$l = -17 \rightarrow 20$

Refinement

Refinement on F^2	Secondary atom site location: difference Fourier map
Least-squares matrix: full	Hydrogen site location: inferred from neighbouring sites
$R[F^2 > 2\sigma(F^2)] = 0.089$	H-atom parameters constrained
$wR(F^2) = 0.172$	$w = 1/[\sigma^2(F_o^2) + (0.0474P)^2 + 0.746P]$
$S = 1.24$	where $P = (F_o^2 + 2F_c^2)/3$
2264 reflections	$(\Delta/\sigma)_{\text{max}} = 0.001$
154 parameters	$\Delta\rho_{\text{max}} = 0.26 \text{ e \AA}^{-3}$
Primary atom site location: structure-invariant direct methods	$\Delta\rho_{\text{min}} = -0.18 \text{ e \AA}^{-3}$
	Extinction correction: none

Special details

Geometry. All e.s.d.'s (except the e.s.d. in the dihedral angle between two l.s. planes) are estimated using the full covariance matrix. The cell e.s.d.'s are taken into account individually in the estimation of e.s.d.'s in distances, angles and torsion angles; correlations

between e.s.d.'s in cell parameters are only used when they are defined by crystal symmetry. An approximate (isotropic) treatment of cell e.s.d.'s is used for estimating e.s.d.'s involving l.s. planes.

Refinement. Refinement of F^2 against ALL reflections. The weighted R -factor wR and goodness of fit S are based on F^2 , conventional R -factors R are based on F , with F set to zero for negative F^2 . The threshold expression of $F^2 > \sigma(F^2)$ is used only for calculating R -factors(gt) *etc.* and is not relevant to the choice of reflections for refinement. R -factors based on F^2 are statistically about twice as large as those based on F , and R -factors based on ALL data will be even larger.

Fractional atomic coordinates and isotropic or equivalent isotropic displacement parameters (\AA^2)

	x	y	z	$U_{\text{iso}}^*/U_{\text{eq}}$
S1	0.50901 (8)	0.0880 (3)	0.62465 (7)	0.0818 (5)
O1	0.2035 (2)	0.1002 (7)	0.53998 (17)	0.0725 (5)
N1	0.3541 (2)	-0.0034 (7)	0.53458 (18)	0.0545 (9)
H1A	0.3909	-0.0822	0.5055	0.065*
C9	0.3780 (3)	0.4046 (9)	0.7161 (2)	0.0610 (12)
H9A	0.3380	0.5543	0.7257	0.073*
H9B	0.4388	0.4763	0.7105	0.073*
C6	0.2391 (3)	-0.2034 (9)	0.4391 (2)	0.0494 (10)
N2	0.3429 (2)	0.2733 (8)	0.64340 (18)	0.0576 (10)
H2A	0.2854	0.2906	0.6265	0.069*
O3	0.4423 (2)	0.3093 (7)	0.84511 (17)	0.0740 (10)
C7	0.2621 (3)	-0.0237 (9)	0.5083 (2)	0.0535 (11)
C1	0.2954 (3)	-0.4110 (9)	0.4191 (2)	0.0555 (11)
H1B	0.3509	-0.4435	0.4501	0.067*
O2	0.3394 (2)	0.0172 (7)	0.7891 (2)	0.0812 (10)
C10	0.3840 (3)	0.2182 (10)	0.7862 (2)	0.0553 (11)
C8	0.3961 (3)	0.1265 (9)	0.6015 (2)	0.0542 (11)
C5	0.1559 (3)	-0.1619 (9)	0.3927 (2)	0.0602 (12)
H5A	0.1168	-0.0250	0.4056	0.072*
C4	0.1308 (3)	-0.3236 (10)	0.3269 (3)	0.0654 (13)
H4A	0.0753	-0.2938	0.2956	0.078*
C3	0.1877 (4)	-0.5258 (11)	0.3085 (3)	0.0704 (14)
H3A	0.1707	-0.6352	0.2647	0.084*
C2	0.2700 (3)	-0.5688 (10)	0.3542 (3)	0.0640 (13)
H2B	0.3088	-0.7062	0.3409	0.077*
C11	0.4497 (4)	0.1530 (13)	0.9180 (3)	0.0918 (19)
H11A	0.4937	0.2362	0.9568	0.138*
H11B	0.4698	-0.0259	0.9074	0.138*
H11C	0.3906	0.1454	0.9381	0.138*

Atomic displacement parameters (\AA^2)

	U^{11}	U^{22}	U^{33}	U^{12}	U^{13}	U^{23}
S1	0.0492 (7)	0.1335 (14)	0.0612 (8)	0.0052 (8)	-0.0012 (5)	-0.0188 (8)
O1	0.0499 (18)	0.093 (3)	0.075 (2)	0.0001 (17)	0.0082 (15)	-0.0248 (19)
N1	0.0436 (19)	0.070 (3)	0.049 (2)	0.0069 (17)	0.0017 (15)	-0.0076 (18)
C9	0.065 (3)	0.059 (3)	0.058 (3)	-0.006 (2)	0.006 (2)	-0.009 (2)

supplementary materials

C6	0.049 (2)	0.055 (3)	0.045 (2)	-0.014 (2)	0.0068 (19)	0.003 (2)
N2	0.052 (2)	0.066 (3)	0.054 (2)	0.0030 (19)	-0.0007 (17)	-0.0090 (19)
O3	0.073 (2)	0.095 (3)	0.0531 (18)	-0.0153 (19)	0.0021 (15)	-0.0035 (18)
C7	0.048 (2)	0.058 (3)	0.055 (3)	-0.002 (2)	0.008 (2)	0.002 (2)
C1	0.062 (3)	0.051 (3)	0.054 (3)	-0.004 (2)	0.005 (2)	0.001 (2)
O2	0.092 (2)	0.065 (2)	0.086 (2)	-0.017 (2)	0.0060 (18)	0.0004 (19)
C10	0.051 (3)	0.056 (3)	0.060 (3)	0.001 (2)	0.014 (2)	-0.009 (2)
C8	0.054 (3)	0.062 (3)	0.045 (2)	-0.001 (2)	0.0012 (19)	0.001 (2)
C5	0.058 (3)	0.060 (3)	0.063 (3)	-0.007 (2)	0.008 (2)	-0.003 (2)
C4	0.059 (3)	0.075 (4)	0.059 (3)	-0.016 (3)	-0.006 (2)	0.006 (3)
C3	0.093 (4)	0.063 (3)	0.054 (3)	-0.018 (3)	0.005 (3)	-0.004 (3)
C2	0.080 (3)	0.051 (3)	0.062 (3)	0.000 (2)	0.011 (2)	-0.005 (2)
C11	0.098 (4)	0.123 (5)	0.055 (3)	0.017 (4)	0.008 (3)	0.011 (3)

Geometric parameters (\AA , $^\circ$)

S1—C8	1.661 (4)	O3—C11	1.452 (5)
O1—C7	1.223 (5)	C1—C2	1.368 (6)
N1—C7	1.371 (5)	C1—H1B	0.93
N1—C8	1.387 (5)	O2—C10	1.197 (5)
N1—H1A	0.86	C5—C4	1.389 (6)
C9—N2	1.437 (5)	C5—H5A	0.93
C9—C10	1.499 (6)	C4—C3	1.361 (6)
C9—H9A	0.97	C4—H4A	0.93
C9—H9B	0.97	C3—C2	1.372 (6)
C6—C1	1.384 (6)	C3—H3A	0.93
C6—C5	1.386 (5)	C2—H2B	0.93
C6—C7	1.482 (6)	C11—H11A	0.96
N2—C8	1.326 (5)	C11—H11B	0.96
N2—H2A	0.86	C11—H11C	0.96
O3—C10	1.318 (5)		
C7—N1—C8	129.0 (3)	O2—C10—C9	124.3 (4)
C7—N1—H1A	115.5	O3—C10—C9	111.3 (4)
C8—N1—H1A	115.5	N2—C8—N1	117.4 (4)
N2—C9—C10	112.4 (4)	N2—C8—S1	124.0 (3)
N2—C9—H9A	109.1	N1—C8—S1	118.6 (3)
C10—C9—H9A	109.1	C6—C5—C4	120.4 (4)
N2—C9—H9B	109.1	C6—C5—H5A	119.8
C10—C9—H9B	109.1	C4—C5—H5A	119.8
H9A—C9—H9B	107.9	C3—C4—C5	119.8 (4)
C1—C6—C5	118.6 (4)	C3—C4—H4A	120.1
C1—C6—C7	123.6 (4)	C5—C4—H4A	120.1
C5—C6—C7	117.9 (4)	C4—C3—C2	120.2 (5)
C8—N2—C9	122.3 (4)	C4—C3—H3A	119.9
C8—N2—H2A	118.9	C2—C3—H3A	119.9
C9—N2—H2A	118.9	C1—C2—C3	120.5 (5)
C10—O3—C11	116.2 (4)	C1—C2—H2B	119.8
O1—C7—N1	121.8 (4)	C3—C2—H2B	119.8
O1—C7—C6	122.7 (4)	O3—C11—H11A	109.5

N1—C7—C6	115.5 (4)	O3—C11—H11B	109.5
C2—C1—C6	120.5 (4)	H11A—C11—H11B	109.5
C2—C1—H1B	119.7	O3—C11—H11C	109.5
C6—C1—H1B	119.7	H11A—C11—H11C	109.5
O2—C10—O3	124.4 (4)	H11B—C11—H11C	109.5

Hydrogen-bond geometry (\AA , $^\circ$)

<i>D</i> —H \cdots <i>A</i>	<i>D</i> —H	H \cdots <i>A</i>	<i>D</i> \cdots <i>A</i>	<i>D</i> —H \cdots <i>A</i>
N2—H2A \cdots O1	0.86	2.02	2.676 (4)	132
C9—H9B \cdots S1	0.97	2.68	3.027 (4)	101
N1—H1A \cdots S1 ⁱ	0.86	2.77	3.547 (3)	151
C9—H9A \cdots O2 ⁱⁱ	0.97	2.54	3.358 (6)	142
C9—H9A \cdots O2 ⁱⁱⁱ	0.97	2.58	3.211 (5)	123

Symmetry codes: (i) $-x+1, -y, -z+1$; (ii) $x, y+1, z$; (iii) $-x+1/2, y+1/2, -z+3/2$.

Fig. 1

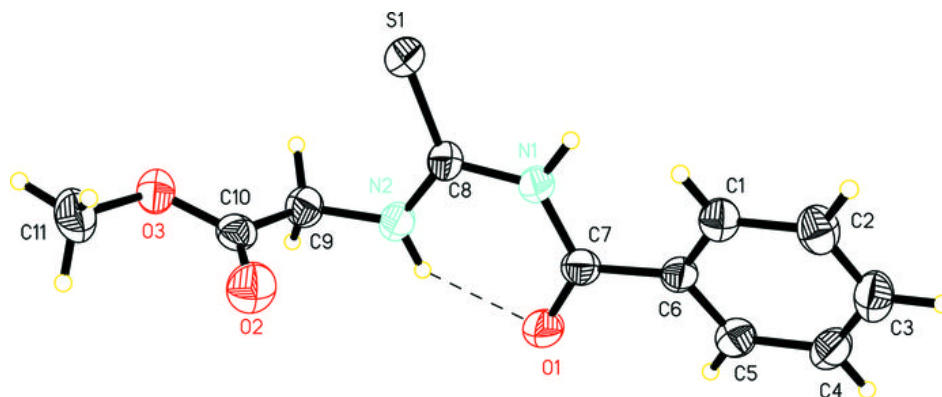


Fig. 2

